

A Novel Method for the Deprotection of N-Boc-D-Glucosamine-Ac by Irradiation via Microwave

H. Bouleghlem, N. E. Aouf, and S. Zidane

Abstract—Since the microwave was used for the first time in 1986 in the field of the organic synthesis, the reactions assisted by microwaves are used more and more in the strategies of synthesis. It meets an interest growing because of its great effectiveness and its extreme profitability due to reduced reaction times, better outputs and obtaining purer products. Its particular use, was the subject of a method around whose developed a green chemistry. Several developed, he was devoted, and the last in date is a work C.O. Kappe in 2005. Within the framework of this work, we will try to develop an aspect of the chemistry of the protective groupings: chimioselective protection deprotection of N-Boc-Glucosamine-Acyle by irradiation under domestic microwave. The usual spectroscopic methods (^1H NMR, IR and MS) were made profitable for elucidation synthesized structures.

Index Terms—Chimioselective, deprotection, glucosamine, irradiation, microwave, protection.

I. INTRODUCTION

Heterocycles are a class of compounds in which one or more carbon atom is replaced by a heteroatom such as oxygen, nitrogen, phosphorus, sulfur ... etc. The most common hetero rings containing nitrogen or oxygen.

Approximately two thirds of publications in chemistry concern heterocycles. A very large number of natural substances and therefore drugs are based of heterocyclic.

The heterocyclic's forms the basic skeleton for a wide variety of chemical compounds interest, biological, pharmacological and industrial [1], [2]. Note that two-thirds of the organic compounds known from the literature are heterocyclic [3], such as sugars which are organic molecules very widespread in the major cell types described: bacteria, and in the connective tissues and cell membranes of animals, vegetable cell.

These are compounds formed during photosynthesis. They are very important in the biological world. They are divided mainly of sugars or monosaccharides and osides (see Fig. 1).

Our interest is charged to aminomonosaccharides compounds (amino sugar). We conducted our research on (+)-D-glucosamine hydrochloride, the (+) - D- $\text{GlcNH}_2 \cdot \text{HCl}$, is very known for its physiological role [4], [5].

Glucosamine is a carbohydrate of the family of osamines whose structure is based on glucose.

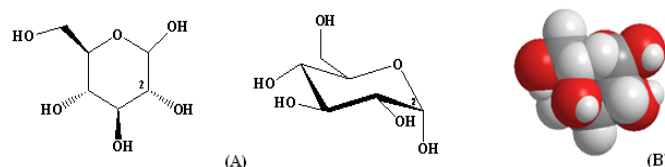


Fig. 1. (+) - D-glucopyranose : representations (A) and hexagonal (B) 3D (red : oxygen; black : carbon ; white : hydrogen) (Space Filling).

In the glucosamine molecule, the carbon at position 2 does not bear alcohol function but a primary amine function ($-\text{NH}_2$) oriented in the same way on the asymmetric carbon (see Fig. 2).

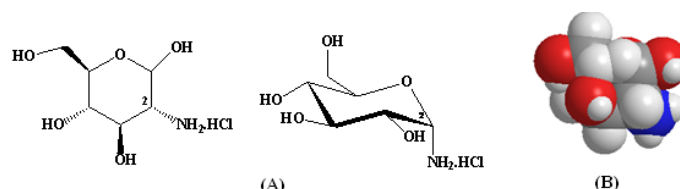


Fig. 2. D - (+) - Glucosamine representations: (A) hexagonal, (B) 3D (red : oxygen; black : carbon ; white : hydrogen; blue : nitrogen) (Space Filling).

Glucosamine is a glycosylated amine. This is a normal constituent of cartilage that would act by stimulating and normalizing the biosynthesis of proteoglycans of the articular cartilage , inhibiting certain enzymes that destroy cartilage and reducing superoxide radicals that damage tissues [6], glucosamine is a precursor used for the glycosylation of proteins and lipids ... etc .

There are two forms of glucosamine differ from their substituents which are: glucosamine hydrochloride and glucosamine sulfate (see Fig. 3) [7].

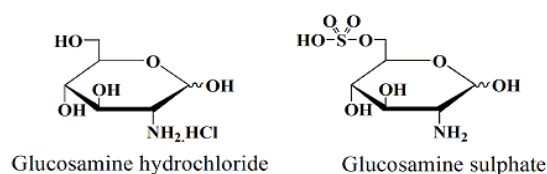


Fig. 3. Representation of (+)-D-Glucosamine-HCl and (+)-D-Glucosamine- SO_3H .

Chemical modification of Glucosamine offers both enormous opportunities and real challenges for chemists. Modifying Glucosamine can indeed help to improve their physicochemical and pharmacological properties.

It is also a way to create original molecules with specific functions, finding various applications beyond the borders of classical chemistry. The challenge facing chemists is to develop effective methods to selectively protect the four alcohol functions and the amine function carried by the glucosamine.

The primary amines are more reactive and less stable because the amino group has one electron pair, which may be

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protonated or reacted with electrophiles.

To make them less reactive the amine can be converted to an amide in the presence of *t*-butoxycarbonyl (Boc).

The cleavage NH-Boc can be achieved for example by using a strong acid $\text{CF}_3\text{CO}_2\text{H}$ (TFA) [8], [9] or by using a chemically inert catalyst such as kaolin or Clay (KC) [10] requires a lot of time with a very low yields by against the cleavage by microwave gives good yields in very short times.

The irradiation of the reactional mediums by the microwaves is a technique increasingly used in organic syntheses. The first organic syntheses carried out in a domestic furnace were described in 1986 [11] by Gedye and Giguere. Today, the number of publications in this field does not cease increasing exponentially [12], [13]. Several authors work in absence of solvent.

The interest to operate without solvent is known for a long time, it relates to the reactivity and sometimes the selectivity of the reactions. The principle is based on the impregnation of the reagents on supports or a simple mixture of reagents with a solid, liquid catalyst or of transfer of phase. The coupling with microwaves makes it possible to operate in full safety with atmospheric pressure in all types of furnaces [14].

The heat transfer under heating microwaves is completely reversed compared to the conventional heating. The traditional heat transfer is transmitted of the outside to the interior of the recipient. Under heating microwave, treated volume becomes him even source of heat. We speak about release of heat for the interior to the outside of the recipient. The external wall of the reactor is colder than the medium of the reactor in the case of the heating microwave, and conversely for the case of the conventional heating by dual envelope, hotplate and flame. It is a mode of instantaneous heating in volume and not on the surface. The thermal phenomena of conduction and convection play a secondary part of equilibrium for the temperature. Local over heating can also occur (see Fig. 4) [15].

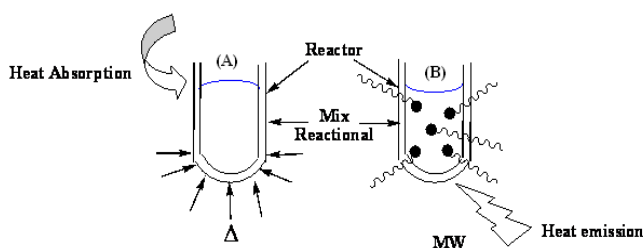


Fig. 4. Gradients in temperature for materials submitted either to traditional (A), or to microwave radiation (B).

Instantaneous character of the heating led to reactions much faster and a weak degradation of the products. Many chemical syntheses use solvents organic, expensive, toxic, non which can be recycled and thus generators of pollution. Thus, the benzene is interdict of employment on an industrial scale, and the use sometimes, of strong mineral acids poses problems of safety. Lastly, the dilution of the reagents lowers the yield and increases considerably the reaction time [16]. (Yield: fraction of reagents having reacted compared to the concerned reagents).

To free itself from solvents thus constitutes an essential objective. This is why the techniques of reaction without solvent developed. It can be a simple mixture of reagents without catalyst [17].

The reaction can proceed if one of the reagents is liquid and if the other reagents are solid, they are soluble in this liquid. Other reactions can, in certain cases, being realized gratitude to the use of mineral supports solid not toxic and which can be recycled like alumina, the silica gels or clay [18], [19].

With the molecular scale, the most polished surface even of a solid is irregular and present cavities. The liquid penetrates in these cavities (adsorption). Its reactivity is strongly increased and the reaction is started with the liquid interface/wall of the cavity. Lastly, a third technique requires the use of inexpensive catalysts making it possible the reaction to always be done without solvent. It is the catalysis by transfer of phase (CTP) solid-liquid without solvent [20], [21].

II. EXPERIMENTAL SECTION

A. Generals Conditions

1) Products

KBr (IR), 98% Acetic Anhydride, 98% KOH, 37% HCl, (+)-D-Glucosamine hydrochloride ((+)-D-GlucNH₂.HCl), from Aldrich Chemical Company, Ethyl acetate (AOEt), Tert-butylidicarbonyl (Boc)₂O, silica gel, sulfate sodium (Na₂SO₄), sodium bicarbonate (Na₂CO₃)

2) Solvents

Methanol 99 % and/ or Ethanol 99.8 %, 99.9% Dichloromethane, Dioxane 99.78 %, distilled water, Hexane, Citric acid, Acetone, Pyridine, DMAP (4-dimethylaminopyridine).

B. Characterization Methods

TLC: The thin layer chromatography was performed on silica plates Merck 60 F354 (Art. 5554).

NMR: The NMR spectra were recorded with a Bruker (Forward) to 400.13 MHz, for ¹H NMR, and 100.61 MHz for ¹³C NMR.

The tetraméthylsilane (TMS) is used as internal reference. Chemical shifts for spectra in D₂O, CDCl₃ or DMSO-*d*⁶ are given in ppm.

Mass spectrometry (MS): Measurements of fast atomic bombardment - mass spectrometry (Matrix: NBA, FABMS) are performed with Auto Spec M (Micromass). The samples were prepared by solution in Meta Nitrophenol.

Infrared Fourier Transform: The FTIR spectra were obtained with a spectrometer Bruker IFS 66/5 from KBr pellets.

In this work, we developed the deprotection of glucosamine (NH-Boc Acyle glucosamine (*a*₁)) by irradiation microwave to form NH₂-Glucosamine Acyle (*a*₂). Total protection was carried out in two stages starting from commercial glucosamine: The protection of the function amine is carried out by Di-tert-Butyl carbonates (Boc)₂O in basic medium, this stage makes it possible to release the amine and to graft Boc, we reaches easily (the NH-Boc under microwave, the chimioselective deprotection and we reaches NH₂-Glucosamine tetra-acetyl (*a*₂) easily. All the reactions are carried out under argon. Their advance is controlled by TLC and the revelation by a solution of H₂SO₄ with (10%), sugars appear in the form of brown spot on TLC.

The purifications are carried out on column of

chromatography on silica gel (35 μm - 70 μm). The purity of the products is controlled by analysis ^1H NMR, IR, by mass spectrometry in chemical mode ionization ESI and by the measurement of the melting point by Electrothermale (BI).

Stage 1: The protection of the function amine by (Boc) $_2\text{O}$ Synthesis of 2-Desoxy-2-N-Tertiobutyloxycarbonyl Glucose (a_1')

In a balloon of 250 mL, 1eq of 2-Desoxy-2-amino-glucose (10g; 49,7 mmol) is solubilized in 150 mL of a mixture of Dioxane and a solution of Potassium hydroxide 1N (1/1) (Dioxane/KOH). To the solution is added 2 eq of (Boc) $_2\text{O}$ (15, 2g; 99,5mmol), the solution is left under agitation at room temperature. At the end of three hours of agitation, a precipitate was formed, this one is filtered then washed with a mixture of water /dioxane, the product of the reaction is obtained in the form of a white powder (see Fig. 5).

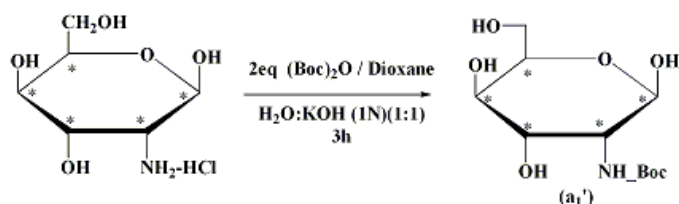


Fig. 5. The protection of the function amine by (Boc) $_2\text{O}$.

Results and Discussion:

The recrystallization is carried out by mixture of three solvent (MeOH / Dioxane / H $_2\text{O}$) (50:25:25) until the formation of the beautiful white crystals, their melting points $T = 194\text{ }^\circ\text{C}$ - $195\text{ }^\circ\text{C}$, 95% yield, TLC ($R_f = 0,2$ [AOEt / H $_2\text{Cl}_2$: v/v ; 8:2] ; $[\alpha]_D^{25} = +64$ ($C = 4,7$; MeOH). ^1H NMR (DMSO, d^6 , 250MHz) δ (ppm) ; 4,96 (d, 1H, H1, $J_{H1-H2} = 3,4\text{Hz}$) ; 3,4- 3,65 (m, 4H, H3, H6, H6') ; 3 - 3,4 (m, 2H, H2, H5) ; 1,35 (s, 9H, Boc).

Mass (Matrix: GT ; FAB $^+$) : 280 [M+H] $^+$; 206 ([M - OC(CH $_3$) $_3$] $^+$).

Stage 2: The protection of all alcohols functions by acetic Anhydride

1) Synthesis of

2-Desoxy-2-N-Tertiobutyloxycarbonyl-2-3-4-6-Tetra-Acetylglucose (a_1'')

1eq of made up NH-Boc-Glucosamine (2.5g, 9.43 Mmol) is solubilized in 10mL of pyridine, a catalytic quantity of 4-Dimethylaminopyridine (DMAP) is added to the solution, and this one is cooled with 0 $^\circ\text{C}$. Then added, drop-by-drop 4eq of acetic Anhydride (7.6mL, 0.075 Mmol) diluted in pyridine 5mL.

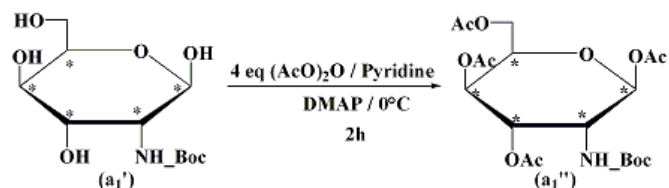


Fig. 6. Protection of all alcohols functions by acetic Anhydride.

At the end of three hours of agitation at ambient temperature, the reactional mixture is diluted in 200 mL of one aqueous solution frozen of NaHO $_3$ with (3%) ; then the phase aqueous and extracted three times with 100 mL from CH $_2\text{Cl}_2$, the

organic phase are gathered, washed with a citric solution of acid (5%), then with water, the organic phase is then dried by sodium sulphate Na $_2\text{SO}_4$, solvent is evaporated under vacuum, the reactional mixture is then purified by chromatography on silica gel column whereas using like leant the system ([AOEt / Hexane : V/V ; 3:7]). 98% yield (see Fig. 6)

Results and Discussion:

TLC ($R_f = 0,5$ [AOEt / Hexane : v/v ; 1:1]). $F = 67\text{ }^\circ\text{C}$ - $70\text{ }^\circ\text{C}$. $[\alpha]_D^{25} = +67$ ($C = 0,15$; HCl $_3$). ^1H RMN (DCl $_3$, 250MHz) δ (ppm) ; 6,12 (d, 1H, H1, $J_{H1-H2} = 3,48\text{Hz}$) ; 5,05 - 5,17 (m, 2H, H3 and H4) ; 4,2 (dd, 1H, H6, $J_{H6-H6'} = 4\text{Hz}$, $J_{H6-H5} = 12,42\text{Hz}$) ; 3,58 - 4,25 (m, 3H, H2, H5 and H6) ; 1,95 - 2,12 (s, 12H, AcO $_2$). ^{13}C RMN (DCl $_3$, 200MHz) δ (ppm) ; 91,10 (C1) ; 71,05 (C3) ; 69,72 (C5) ; 67,82 (C4) ; 61,67 (C6) ; 52,26 (C2) ; 28,21 (C(CH $_3$) $_3$) ; 21,58 - 21,93 (COCH $_3$) ; 151,25 (C(CH $_3$) $_3$) ; 154,90 (NHCO) ; 168,70 - 171,13 (COCH $_3$). Masse (Matrix : NBA ; FAB $^+$) : 348 [M + H] $^+$; 370 [M + Na] $^+$; 289 [M - AC - OH - H] $^+$; 246 [M - AC - OH - Ac].

Stage 3: The deportation by the use of Microwave new method use the modern heating by magnetic waves "Micro-Wave" (MW)

In a domestic Microwave of (70 W to 600 W) LG marks, in a erylen 50mL poses 0.5g NH-Boc-Glucosamine-Ac, we adds mineral solid supports, the reactional mixture is left during 1 min with 350 Watt until the end of the reaction, evolution of the reaction is followed by TLC, which shows the polar formation of one produced revealed in Ninhydrine (see Fig. 7).

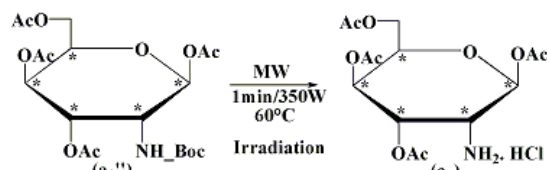


Fig. 7. Deportation by the use of microwave.

Results and Discussion:

The precipitated product is filtered by ether to lead a white powder in 90% yield. TLC: $R_f = 0,4$ [AOEt / H $_2\text{Cl}_2$. v/v ; 8:2]. ^1H NMR (DMSO, d^6 , 250MHz) δ (ppm) ; 6,2 (d, 1H, H1, $J_{H1-H2} = 3,62\text{Hz}$) ; 5,25(dd, 1H, H3 , $J_{H2-H3} = 9,89\text{Hz}$, $J_{H3-H4} = 9,73\text{Hz}$) ; 4,92 (dd, 1H, H4, $J_{H4-H5} = 9,38\text{Hz}$) ; 4,12 - 4,20 (m, 2H, H6, H6') ; (3,5 - 3,85) (m, 2H,H5,H2). ^{13}C NMR (DMSO, d^6 , 200MHz) δ (ppm) ; 89,01 (C1) ; 71,02 (C3) ; 70,59 (C5) ; 69,72 (C4) ; 61,87 (C6) ; 51,01 (C2) ; 21,16-21,94 (COCH $_3$) ; 169,53 - 170,84 (COCH $_3$).

TABLE I: COMPARISON BETWEEN THE DEPROTECTION BY MICROWAVES, DEPROTECTION BY THE TFA AND DEPROTECTION BY CATALYZE

	Time	Reagent Quantity	Yield %
Deportation by MW	1min to 350W	---	90
Deportation by TFA	5h with 0 $^\circ\text{C}$	20 mL of TFA	25
Deportation by Catalyze	8h with 100 $^\circ\text{C}$	5g of KC	40

The results are summarized in Table I. The reaction is quantitative under microwaves in 1 minute with power of 350W, whereas the reaction time with the use of the TFA (Trifluoroacetic Acid) and kaolin of Clay (KC) is very important.

III. CONCLUSION

At this contribution on the interest of the microwaves in organic synthesis, we understand better why electrothermic microwave is a technology in full rise. Obviously, this technology fulfills perfectly well the requirements of today which are the speed and the effectiveness. The deprotection by the use of microwave aims to the modernization and the simplification of the processes and to make them less polluting.

ACKNOWLEDGEMENT

The compounds obtained with good yields can be used as precursors in the design of bioactive entities and others by addition or nucleophilic substitution to lead to the analogues glucosamine models. This work can be supplemented by a toxicological study, kinetics of decomposition in physiological medium and others evaluations pharmacological in vivo and in vitro.

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